

PROCESS FOR THE PRODUCTION OF METAL SALTS
OF RADICAL-POLYMERIZABLE COMPOUNDS

0
[0001] The present invention relates to a novel method for manufacturing metal salts of
5 radical-polymerizable compounds with low content of moisture.

BACKGROUND

[0002] Normally, sodium methacrylate, potassium methacrylate, sodium acrylate, potassium
acrylate, or similar metal salts of radical-polymerizable compounds, are obtained by
neutralizing aqueous solutions of alkali-metal compounds. However, the problem associated
10 with the aforementioned process consists in that it is extremely difficult to produce metal salts
of the radical-polymerizable compounds with low moisture content when they are obtained
from aqueous compounds. Many suggestions have been made heretofore to overcome this
problem of the prior-art. For example, there were attempts to remove water by subjecting
water solutions of radical-polymerizable compounds to azeotropic dehydration in the presence
15 of an organic solvent such as toluene. This method, however, cannot produce metal salts of
radical-polymerizable compounds of high purity because the process is accompanied by a
polymerization reaction that occurs during azeotropic dehydration. Attempts have been made
to suppress this polymerization reaction by adding conventional polymerization inhibitors
normally used in conjunction with radical-polymerizable compounds, such as hydroquinone,
20 methoxyhydroquinone, phenothiazine, hindered phenol, or the like. However, since all these
compounds are oil-soluble, their solubility in water is low, they cannot completely stop
polymerization of radical-polymerizable compounds, and, therefore, are unsuitable for
obtaining metal salts of radical-polymerizable compounds of high purity.

[0003] On the other hand, known in the art is a method for obtaining moisture-free
25 metal salts of a methacrylic acid, wherein an aqueous component is removed by
pulverizing an aqueous solution of a methacrylic acid metal salt in a hot-air circulation
drier (see Japanese Laid-Open Patent Application Publication (hereinafter referred to as
Kokai") 47-031924 and Kokai 49-011820). However, this method is expensive as it
requires the use of specific equipment.

30 [0004] Furthermore, a method is known for manufacturing a water-free potassium

methacrylate obtained by neutralizing a methacrylic acid in a methanol solution of a potassium methyllate (see Kokai 3-209388). However, the potassium methyllate used in this process is an expensive reagent, which is not suitable for realization of low-cost production plans.

5

THE INVENTION

[0005] Based on the study aimed at a solution of the problems of the prior art, the authors arrived at the present invention. It is an object of the present invention to provide a method for high-yield production of high-purity metal salts of radical-polymerizable compounds having a reduced content of aqueous components.

10 [0006] More specifically, the present invention relates to a method for manufacturing metal salts of radical-polymerizable compounds with reduced amount of aqueous components by subjecting an aqueous solution of an alkali-metal salt, alkali earth-metal salt, or a zinc salt of a radical-polymerizable compound (A) selected from the group consisting of a compound represented by the following general formula (1): $R^1-CH=CR^2-COOH$ (where R^1 designates a hydrogen atom, an alkyl group or a phenyl group, and R^2 designates a methyl group or a hydrogen atom), a compound represented by the following general formula (2): $R^1-CH=CR^2-R^3-COOH$ (where R^1 and R^2 are the same as defined above, and R^3 designates a phenylene group), a compound represented by the following general formula(3): $R^1-CH=CR^2-SO_3H$ (where R^1 and R^2 are the same as defined above),
15 and a compound of the following general formula(4): $R^1-CH=CR^2-R^3-SO_3H$ (where R^1 , R^2 , and R^3 are the same as defined above)to heating under a reduced pressure in the presence of (B) an aqueous polymerization inhibitor selected from L-ascorbic acid, sodium L-ascorbate; erythorbic acid sodium erythorbate; N-phenyl hydroxylamine, aqueous salts of N-phenyl hydroxylamine; 3,5-di-t-butyl-4-hydroxyaniline, aqueous salts of 3,5-di-t-butyl-4-hydroxyaniline; 3,5-di-t-butyl-4-hydroxybenzoic acid, and aqueous salts of 3,5-di-t-butyl-4-hydroxybenzoic acid.
20
25

The invention also relates to a method for manufacturing metal salts of radical-polymerizable compounds with reduced amount of aqueous components by subjecting an

aqueous solution of an alkali-metal salt, alkali earth-metal salt, or a zinc salt of a radical-polymerizable compound (A) selected from the group consisting of a compound represented by the following general formula (1): $R^1-CH=CR^2-COOH$ (where R^1 designates a hydrogen atom, an alkyl group, or a phenyl group, and R^2 designates a methyl group or a hydrogen atom), a compound represented by the following general formula (2): $R^1-CH=CR^2-R^3-COOH$ (where R^1 and R^2 are the same as defined above, and R^3 designates a phenylene group), a compound represented by the following general formula (3): $R^1-CH=CR^2-SO_3H$ (where R^1 and R^2 are the same as defined above), and a compound of the following general formula (4): $R^1-CH=CR^2-R^3-SO_3H$ (where R^1 , R^2 , and R^3 are the same as defined above) to heating under a reduced pressure in the presence of (B) an aqueous polymerization inhibitor selected from L-ascorbic acid, sodium L-ascorbate; erythorbic acid sodium erythorbate; N-phenyl hydroxylamine, aqueous salts of N-phenyl hydroxylamine; 3,5-di-t-butyl-4-hydroxyaniline, aqueous salts of 3,5-di-t-butyl-4-hydroxyaniline; 3,5-di-t-butyl-4-hydroxybenzoic acid, and aqueous salts of 3,5-di-t-butyl-4-hydroxybenzoic acid.

[0007] Thus, the method of the present invention makes it possible to reduce the content of aqueous components and to increase the production yield of high-purity salts, such as alkali-metal salts, alkali-earth salts, or zinc salts of radical-polymerizable compounds.

Furthermore, the method prevents occurrence of radical polymerization during the manufacturing process.

Best Mode for Carrying Out the Invention

[0008] A radical-polymerizable compound that constitutes aforementioned component (A) is selected from the group consisting of a compound represented by the following general formula(1): $R^1-CH=CR^2-COOH$ (where R^1 designates a hydrogen atom, an alkyl group, or a phenyl group, and R^2 designates a methyl group or a hydrogen atom), a compound represented by the following general formula (2): $R^1-CH=CR^2-R^3-COOH$ (where R^1 and R^2 are the same as defined above, and R^3 designates a phenylene group), a compound represented by the following general formula(3): $R^1-CH=CR^2-SO_3H$ (where R^1 and R^2 are the same as defined above), and a compound of the following general formula(4): R^1-

$\text{CH}=\text{CR}^2\text{-R}^3\text{-SO}_3\text{H}$ (where R^1 , R^2 , and R^3 are the same as defined above). In the above formulae, R^1 may designate a hydrogen atom, methyl group, ethyl group, propyl group, or a similar alkyl group, of which the hydrogen atom or phenyl group, and in particular hydrogen atom, is most preferable. R^2 may designate a methyl group or a hydrogen atom,
5 and R^3 may designate a phenylene group.

[0009] The aforementioned compounds can be exemplified by a methacrylic acid, acrylic acid, cinnamic acid, o-vinylbenzene-sulfonic acid, p-vinylbenzene-sulfonic acid, p-vinylbenzoic acid, or their derivatives.

[0010] Component (A) may also comprise an aqueous solution of an alkali metal salt,
10 alkali-earth metal salt, or a zinc salt of the aforementioned radical-polymerizable compounds. The alkali metal of such salt may comprise lithium, sodium, potassium, or rubidium. Of these, the most preferable are sodium and potassium. Alkali-earth metals can be represented by magnesium, calcium, strontium, and barium. Of these, most preferable are magnesium and calcium. The following are examples of metal salts of the
15 aforementioned radical-polymerizable compounds: sodium methacrylate, potassium methacrylate, calcium methacrylate, magnesium methacrylate, zinc methacrylate, sodium acrylate, potassium acrylate, magnesium acrylate, calcium acrylate, magnesium acrylate, zinc acrylate, sodium cinnamate, calcium cinnamate, sodium p-vinyl benzoate, potassium p-vinyl benzoate, sodium o-vinylbenzene sulfonate, potassium o-vinylbenzene sulfonate,
20 sodium p-vinylbenzene sulfonate, and potassium p-vinylbenzene sulfonate.

[0011] Aqueous solutions of metal salts of the aforementioned radical-polymerizable compounds can be easily prepared by conventional methods, e.g., by mixing the radical-polymerizable compounds with aqueous solutions of alkali-metal salts, alkali-earth metal salts, or zinc salts and thus neutralizing the compounds, or by saponifying esters of the
25 aforementioned radical-polymerizable compounds with aqueous solutions of the alkali-metal salts, alkali-earth metal salts, or zinc salts.

[0012] An aqueous polymerization inhibitor that constitutes component (B) is a specific component of the present invention. This component can be represented by L-ascorbic

acid, sodium L-ascorbate, erythorbic acid, sodium erythorbate, or derivatives of the above compounds; N-phenylhydroxylamine, aqueous salts of N-phenyl hydroxylamine; 3,5-di-t-butyl-4-hydroxyaniline, aqueous salts of 3,5-di-t-butyl-4-hydroxyaniline; 3,5-di-t-butyl-4-hydroxybenzoic acid; aqueous salts of 3,5-di-t-butyl-4-hydroxybenzoic acid. The most
5 preferable among the above are L-ascorbic acid (vitamin C), N-phenylhydroxylamine, and hydrochloric salts of N-phenylhydroxylamine, in particular L-ascorbic acid (vitamin C). These compounds are almost completely non-toxic and at the same time are extremely efficient inhibitors of polymerization for radical-polymerizable compounds.

[0013] Component (B) should be used in an amount sufficient for inhibiting
10 polymerization of alkali-metal salts, alkali-earth metal salts, or zinc salts of the radical-polymerizable compounds of component (A), but, in general, it is recommended to use them in an amount of 0.001~0.1 parts by weight for each 100 parts by weight of the alkali-metal salts, alkali-earth metal salts, or zinc salts of the radical-polymerizable compounds of component (A). Component (B) can be added at the stage of synthesis of
15 component (A) or after the synthesis of component (A). Addition in the synthesis stage is preferable for the reason of simplicity of the operation.

[0014] In accordance with the present invention, aqueous components are removed when component (A) is heated under a reduced pressure in the presence of aforementioned component (B). The use of conventional reaction equipment for this purpose may cause
20 some problems because removal of water from component (A) increases the content of solids and thus hinders stirring. Therefore, components (A) and (B) are first combined with an organic solvent, such as toluene, xylene, heptane, octane, or the like, and then heated for removal of water from component (A) by distillation. Alternatively, components (A) and (B) are heated under a reduced pressure, and, after a major portion of water is
25 removed from component (A), the product is subjected to azeotropic dehydration. The process can be conducted in a mixer-drier suitable for stirring powdered materials, especially for dehydration without addition of organic solvents.

[0015] The method of the present invention makes it possible to obtain high-purity salts,

such as alkali-metal salts, alkali-earth metal salts, and zinc salts of radical-polymerizable compounds, with high yield and essentially without aqueous components, the content of

which does not exceed 5 wt.%, preferably does not exceed 0.5 wt.%, and even more

5 preferably does not exceed 0.2 wt.%.

[0016] Thus, as has been mentioned above, the alkali-metal salts, alkali-earth metal salts, and zinc salts of radical-polymerizable compounds produced by the method of the present invention are characterized by high purity and low moisture content. Due to these characteristics, the salts of the present invention are suitable for use as intermediate
10 starting materials in reactions with water-unstable organic compounds. They also can be used as additives to such compounds for improving their physical properties.

Examples

[0017] The invention will be further described in more detail with reference to practical examples.

15

Example 1

[0018] A mixture prepared from 284g (3.30 moles) of methacrylic acid, 280g of water, and 500 mg of L-ascorbic acid (vitamin C) was combined with a 48 wt.% aqueous solution of potassium hydroxide and the mixture was stirred to produce an aqueous solution of potassium methacrylate. The obtained aqueous solution was loaded into a mixer-drier and
20 was subjected to dehydration by distillation with heating at 60°C under a reduced pressure of 20mmHg. Measurement showed that content of water in the obtained potassium methacrylate did not exceed 0.1 wt.%. The content of water in the product was determined by weighing the distilled water. The potassium methacrylate was obtained with a yield of 98%. The above reaction did not produce any products of polymerization
25 that could be caused by the presence of methacrylic acid.

Example 2

[0019] A mixture prepared from 284 g (3.30 moles) of methacrylic acid, 280 g of water, and 500 mg of L-ascorbic acid (vitamin C) was combined with a 48 wt.% aqueous solution

of potassium hydroxide and the mixture was stirred to produce an aqueous solution of potassium methacrylate. The obtained aqueous solution was subjected to dehydration by distillation with heating under a reduced pressure of 150 mm Hg. The process was continued until the residual content of water reached 280 g. Following this, the pressure
5 in the reaction system was restored to normal, 600 g of toluene were added, and the product was subjected to azeotropic dehydration to remove the residual moisture and obtain a mixture of potassium methacrylate with toluene. The content of water in potassium methacrylate was 0.03 wt.%. The content of water in the product was determined by means of the Karl Fischer method. The potassium methacrylate was
10 obtained with a yield of 99%. The above reaction did not produce any products of polymerization that could be caused by the presence of methacrylic acid.

Example 3

[0020] A mixture prepared from 284g (3.30 moles) of methacrylic acid, 280g of water, and 3,5-di-t-butyl-4-hydroxyaniline hydrochloride was combined with a 48 wt.% aqueous
15 solution of potassium hydroxide and the mixture was stirred to produce an aqueous solution of potassium methacrylate. The obtained aqueous solution was loaded into a mixer-drier and was subjected to dehydration by distillation with heating at 60 °C under a reduced pressure of 20 mm Hg. Measurement showed that content of water in the obtained potassium methacrylate did not exceed 0.05 wt.%. The content of water in the
20 product was determined by weighing the distilled water. The potassium methacrylate was obtained with a yield of 97%. The above reaction did not produce any products of polymerization that could be caused by the presence of methacrylic acid.

Example 4

[0021] A mixture prepared from 859 g of an aqueous solution of sodium vinylsulfonate
25 (the product of Aldrich Co., 25 wt.% aqueous solution, 1.65 moles of sodium vinylsulfonate) and 300 mg of L-ascorbic acid(vitamin C) was loaded into a mixer-drier and was subjected to dehydration by distillation with heating at 50 °C under a reduced pressure of 10mmHg. Measurement showed that content of water in the obtained sodium vinylsulfonate did not

exceed 0.05 wt.%. The content of water in the product was determined by weighing the distilled water. The sodium vinylsulfonate was obtained with a yield of 96%. The above reaction did not produce any products of polymerization that could be caused by the presence of vinylsulfonic acid.

5

Comparative Example 1

[0022] A mixture was prepared from 284g (3.30 moles) of methacrylic acid, 280g of water, and 500mg of methoxyhydroquinone (MEHQ). The mixture was combined with a 48 wt.% aqueous solution of potassium hydroxide and stirred to produce an aqueous solution of potassium methacrylate. The obtained aqueous solution was subjected to dehydration by distillation with heating under a reduced pressure of 150 mm Hg. The process was carried out until the content of water reached 280 g. Following this, the pressure in the reaction system was restored to normal, 600 g of toluene were added, and the product was subjected to azeotropic dehydration to remove the residual moisture and obtain a potassium methacrylate. The potassium methacrylate was obtained with a yield of 83%.
10
15 The above reaction system produced a viscous product of polymerization that was caused by the presence of methacrylic acid.

Comparative Example 2

[0023] The process for obtaining potassium methacrylate was the same method as in Example 2, with the exception that 2,6-di-*t*-butyl-4-methylphenol (BHT) was used instead of L-ascorbic acid. In the dehydration process, BHT was distilled out together with water, and azeotropic dehydration produced viscous products of polymerization that hindered stirring and did not allow preparation of the target potassium methacrylate.
20